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Comment on “Glyphosate and Fluoride in High-Hardness Drinking Water Are Positively Associated with Chronic Kidney Disease of Unknown Etiology (CKDu) in Sri Lanka”



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Ulrich et al.¹ claim a positive association among glyphosate (GLY), fluoride, high-hardness drinking water (HW), and chronic kidney disease of unknown etiology (CKDu). Invoking claims of glyphosate causing kidney damage and mitochondrial damage in zebrafish, they implicate glyphosate in the etiology of CKDu. In fact, we show below that their data taken as is provide strong evidence that glyphosate in any form is not implicated in the etiology of CKDu.

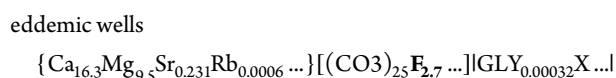
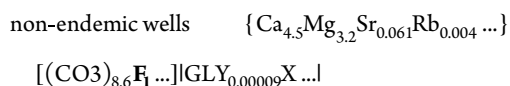
Ulrich et al.¹ report that fluoride (F⁻), hard water, V, Sr, Rb, and GLY are strongly correlated with an odds ratio of >0.5, with glyphosate being nearly 3.5 times more associated with fluoride for endemic wells. Fluoride and glyphosate correlation is their main claim.

The data analysis uses concentrations expressed in grams per liter, etc., and not in moles per liter, thus ignoring differences in molecular and atomic weights (e.g., GLY, molecular weight of 169; Li, atomic weight of 6.9), producing nonphysical correlations that exaggerate GLY by a factor of ~200.

Hard water (HW) is a mixture of hydrated Ca and Mg cations and anions such as CO₃, HCO₃, SO₄, and even PO₄. These need separate treatment if other alkaline earths (such as Sr and Rb) and fluoride are treated as independent variables, especially for the chemistry of regolith aquifers. Fluoride is the only anion treated in ref 1. We also include CO₃ for the sake of clarity. Humic (organic) and silicic acids, bases, and ion solvation are also ignored in ref 1. We denote them as X.

We treated HW explicitly as a standard mixture of Ca and Mg carbonates. Chemical associations are reported in ref 1 in the form {Ca₁₁Mg₁₂Sr₁₃, ..., etc.} for cations and [F₁₁(CO₃)₁₂...] for anions, using molar concentrations. GLY and other chemical moieties of interest are written as |GLY_{k1}X_{k2}...|.

Using the average molar concentrations based on the 75th percentile given in Tables S3–S6 of the Supporting Information of ref 1, we rewrite ionic-association formulas as follows. The number of fluoride ions is normalized to unity for non-endemic wells.



This “molecular” representation of the Ulrich data¹ shows that while there are nine glyphosate moieties per 10 000

fluoride ions in non-endemic well water, there are <12 GLY moieties in endemic wells. Therefore, the difference in the probability of a fluoride ion encountering a glyphosate molecule in an endemic well, as compared to a non-endemic well, is a mere 0.03%. Furthermore, some non-endemic wells with 2.97 μg/L show more glyphosate per fluoride ion. Therefore, the “increased” correlation with fluoride in CKDu-endemic wells has no significance, being within well-to-well random fluctuations.

The difference between endemic and non-endemic wells becomes even smaller if the Ca–glyphosate chelate (molecular weight of 209) is used as the relevant chemical moiety. Rates of chemical reactions as well as toxicokinetic rate processes are proportional to the concentrations of the reacting species or to some power thereof. By this analysis, there is no difference in toxicity significance in non-endemic wells and endemic wells.

The Gibbs free energy of a cluster of ions, inclusive of solvation effects, can indicate ion clustering. Such calculations revealed² that fluoride and magnesium ions may persist as (MgF)⁺-type transients. The synergy of such complexes on CKDu was established by Wasana et al.³ Similar calculations with the GLY component, even at higher concentrations than in CKDu-endemic wells, show no stabilization because the chelation generates very low ambient concentrations of GLY. In effect, hard water protects against any possible glyphosate toxicity but enhances fluoride toxicity.

CKDu may arise from cumulative effects of environmental agents whose nature is mired in controversy.⁴ Fluoride and possibly magnesium ions present together have been favored in several studies.⁵ In the assessment with pesticide effects, a time average taken over at least two planting seasons is necessary. Correlations, when found in a single-shot study while marred by some control wells being in noncomparable climate zones, e.g., at Matara compared to Padaviya, are not meaningful because multiple time measurements are needed.

According to Ulrich et al.,¹ “Glyphosate can tightly chelate divalent metal cations ... These glyphosate–metal complexes increase the half-life of the herbicide dramatically, from 90 days

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to 7 years in water and from 47 days to 22 years in soil. This behavior makes glyphosate a potential contributor to CKDu”.

Such long persistence established for cold climates is sharply reduced to a few days for the Sri Lankan dry zone where the average temperature T that is present in the Arrhenius factor in the decay rate exceeds 30 °C. The tightly chelated glyphosate–metal complex is NOT glyphosate and not a herbicide but a strongly inert insoluble substance. The presence of some algae and aquatic weeds in these wells shows that there is no active herbicide in the water. The dose used is said to be “metal concentrations ... found in CKDu endemic reservoirs and wells”. If the latter is a threshold for humans (average weight of 60 kg), then the dose used for zebrafish (average weight of 150 mg) is some 400 000 times the human threshold.

What has been studied with zebrafish is acute toxicity and has no relevance to chronic toxicity.

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